

Chemistry 3719 - Organic Chemistry I

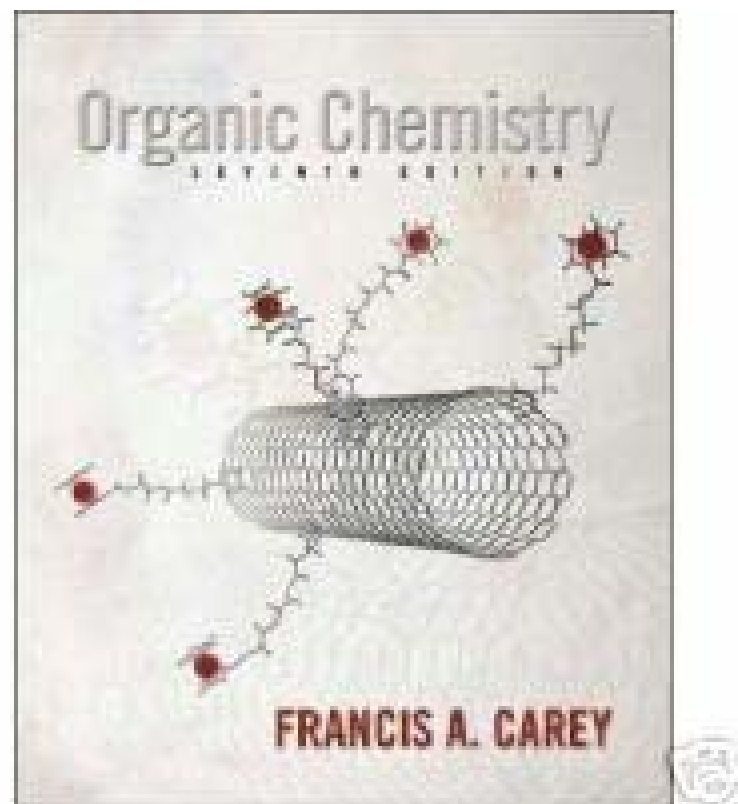
Professor : Dr. Peter Norris
Office : 6014 Ward Beecher
Telephone : (330) 941-1553
Email : pnorris@ysu.edu
Websites :

http://www.as.ysu.edu/~pnorris/public_html

www.chemfinder.com

Lecture needs:

- **Carey**
- **Molecular models**
- **Adobe Acrobat Reader**
- **Web access**



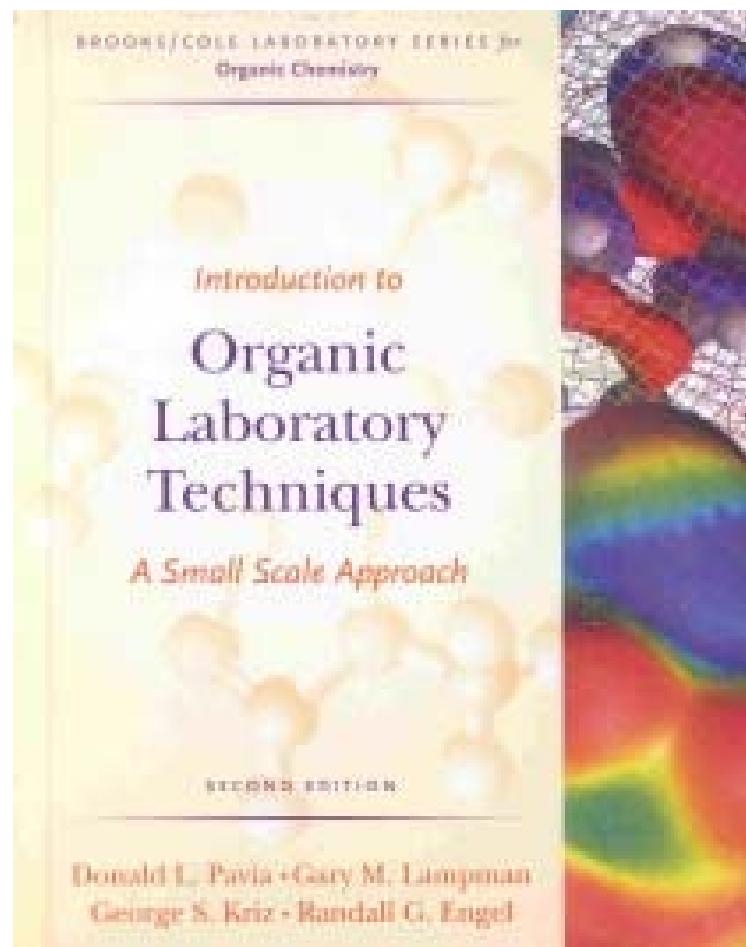
Molecular Models – www.darlingmodels.com



May be used on exams, will be used in lecture

Lab needs:

- **Pavia, Lampman, Kriz and Engel**
- **Goggles**
- **Lab coat**
- **Bound notebook**



Chemistry 3719 and 3720 (and labs)

Lectures

- Structure and nomenclature of compounds and groups
- Physical properties and analysis of materials
- **Reactivity and transformations with reagents**
- Importance of organic compounds in other subjects

Labs

- Glassware and equipment used to prepare organics
- Instrumentation used to analyze compounds
- Keeping a good notebook of lab preparations

Chemistry 3719R and 3720R (Recitation)

Objectives

- Practice the problems sets, old exams
- Practice the problems from the book
- Ask ?? of a professional chemist (other than lecture Prof)
- To encourage students to keep up with material (quizzes)

When: 12-12.50 or 1-1.50 on Mondays

(1 Semester hour, Separate grade to 3719/3719L)

Chemistry 3719 Personnel



Dr. Peter Norris
3719 lecture



Dr. John Jackson
3719 recitation



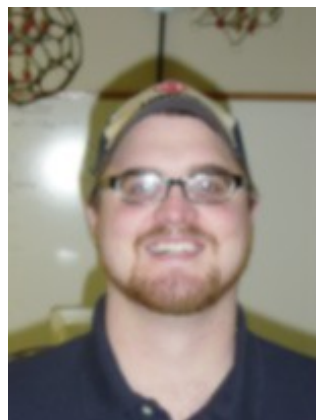
Calvin Austin
3719 lab



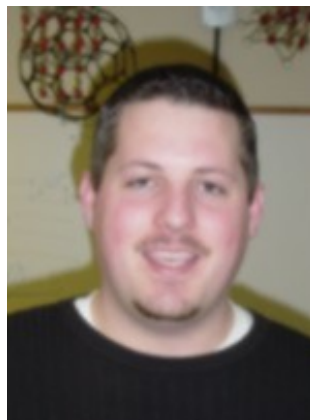
Lucas Beagle
3719 lab



Lemuel Carlisle
3719 lab



Brian Dobosh
3719 lab



Mike Evans
3719 lab



Ashley Malich
3719 lab



Kevin White
3719 lab

Some does and don'ts for 3719 and 3720

Does

- Prepare for lecture and lab; **read ahead**
- Ask questions at any time; lecture, recitation, office hours
- Use all of your resources; email, website, tutors
- If you struggled in General Chemistry, seek help soon

Don'ts

- **Don't get behind**, blow off class, ignore the available help
- Don't wait until October to say “dude, I thought I knew the stuff.”
- Don't complain when you get 20/100 if you ignore the above

Get help : pnorris@ysu.edu

Peter Norris B.Sc., Ph.D.

Born : 1965, Liverpool, England

B.Sc. Chemistry : 1986, Salford University, England

Ph.D. Organic Chemistry: 1992, The Ohio State University

Post-doctoral : 1993-96, American University, Wash'n DC

Assistant Professor : 1996-2000 YSU Chemistry

Associate Professor : 2000-2004 YSU Chemistry

Full Professor : 2004 – present YSU Chemistry

40 publications, graduated 23 Masters degree students since 1998

~ \$1,000,000 in grant money since 1999

Research and Publication

"Crystal and molecular structure of 6,7-dideoxy-1,2;3,4;9,10-tris-O-(1-methylethylidene)-D-erythro-alpha-D-galacto-undecopyranosid-8-ulose," **T. D. Weaver**, M. Zeller, and P. Norris, *J. Chem. Cryst.*, **2006**.

"N-Glycoside neoglycotrimers from 2,3,4,6-tetra-O-acetyl-beta-D-glucopyranosyl azide," **D. P. Temelkoff**, M. Zeller, and P. Norris, *Carbohydrate Research* **2006**, 341, 1081-1090.

"Application of Bis(diphenylphosphino)ethane in Staudinger-type N-Glycosyl Amide Synthesis," **D. P. Temelkoff**, **C. R. Smith**, **D. A. Kibler**, S. McKee, **S. Duncan**, M. Zeller, M. Hunsen, and P. Norris, *Carbohydrate Research*, **2006**, 341, 1645-1656.

"Crystal structure of 1-(2,3:5,6-di-O-isopropylidene-beta-D-mannofuranosyl)-1H-[1,2,3]triazol-4,5-dicarboxylic acid diethyl ester," H. Seibel, **P.L. Miner**, P. Norris, and T.R. Wagner, *J. Chem. Cryst.*, **2006**.

"Cu(I)-Catalyzed formation of D-mannofuranosyl 1,4-disubstituted 1,2,3-triazole carbohybrids," **P.L. Miner**, T.R. Wagner, and P. Norris, *Heterocycles* **2005**, 65, 1035-1049.


>40 total, most with YSU undergrad or MS students as coauthors

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http://www.as.ysu.edu/~pnorris/

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Peter Norris, Ph.D.

Organic Chemistry

Professor
Youngstown State University, 2004-present

Associate Professor
Youngstown State University, 2000-2004

Assistant Professor
Youngstown State University, 1996-2000

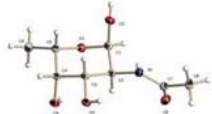
Post-doctoral
American University, Washington, DC, 1993-1996

Ph.D., Organic Chemistry
The Ohio State University, Columbus, OH, 1992

B.Sc. (Hons.), Chemistry
Salford University, Salford, England, 1986

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- ▶ Chemistry 1516L
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- ▶ Chemistry 4850
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- ▶ Chemistry 6941
- ▶ Chemistry 6942
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- ▶ Grad School Info
- ▶ NMR Practice
- ▶ NMR Unknowns
- ▶ Structure Unknowns
- ▶ Monosaccharide NMR

We have a general interest in the synthesis of small molecules to serve as inhibitors of biological processes. Group projects are currently funded by the National Institutes of Health (AREA) and the American Chemical Society (Petroleum Research Fund).



The X-Ray crystal structure of bacterial aminosugar *N*-acetyl-L-fucosamine recently solved at YSU (A-B Alhassan, M Zeller and P. Norris **2006**).

Contact:

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[Youngstown State University](#)
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Youngstown, OH 44555

6014 Ward Beecher Hall
Phone: 330-941-1553
Fax: 330-941-1579
E-mail: pnorris@ysu.edu

News:

Ryan Conway (2nd year MS student) has accepted a position in Process Chemistry with [Roche Pharmaceuticals](#) in Florence, SC, and **Adam Cox** (2nd year MS student) is moving to the Ph.D. program in Organic Chemistry at [The Ohio State University](#). Both will graduate with their MS degrees in August 2007.

Graduate students from the Norris group at YSU presented posters on their work at the [2007 Central Regional Meeting](#) of the American Chemical Society, held in Covington, Kentucky May 20-23 2007.

Professor Diana Fagan, YSU Department of Biological Sciences, presented the paper [Carbohydrate Mimetics for *S. aureus* capsule](#) at the [2007 Experimental Biology](#) meeting held in Washington, DC, April 28 to May 2 2007 ([Link to abstract](#)). The work represents initial biological assessment of some of the glycomimetic molecules created in the Norris lab at YSU.

NSF awards \$475,000 for Chemistry - the YSU Department of Chemistry was awarded \$475,000 from the National Science Foundation for the purchase of a new remotely accessible 400 MHz Nuclear Magnetic Resonance Spectrometer (CRIF:MU: CAMRA-YSU - Cyber-Accessible Magnetic Resonance at YSU - Purchase of a 400 MHz NMR Spectrometer). The new NMR will be installed in 2007 ([Link to article](#)).

**Youngstown State University
Chem 3719 - Organic Chemistry I**

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Old Exams

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Exam Keys

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[Fall 2006 Exam 3 key](#)

[Fall 2005 Exam 3 key](#)

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[Acid-Base Answers](#)

[Stereochem Answers](#)

pnorris@ysu.edu

YSU

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Who	When	Messages
Post a new message		
nina	413 05-14-2007 04:43 PM ET (US)	I just wanted to let you know that I learned more ochem I in your ochem II class than I did in my ochem I class---highly appreciated:)
Peter Norris	412 05-09-2007 07:54 PM ET (US)	Ashley - email me directly at pnorris@ysu.edu.
Ashley Whitfield	411 05-09-2007 07:49 PM ET (US)	Dr. Norris, Can i have my grade please? Thank you. QT - Peter Norris wrote: > < replied-to message removed by QT >
Peter Norris	410 05-09-2007 05:20 PM ET (US)	Grades are done. Email me at pnorris@ysu.edu for the details. 134/200 average on the final. High of 196/200 (scary...), low of 34/200. I had a lot of fun this year with 3719/3720, I enjoyed seeing people rise to the challenge and "get stuck in" as we would say in the UK. The coming changes at YSU are a real chance to get rid of "You Screwed Up" and replace it with something positive (okay, I'm still working on this) like "You Succeeded, Unequivocally" (I'd make a lousy marketing major). I'm optimistic about YSU, which I haven't really been before. Enjoy the glorious weather and all the best for the future. Peter Norris
Jamie	409 05-07-2007 07:58 AM ET (US)	and also, when counting H neighbors, do would you stop counting carbons at a carbonyl?
Jamie	408 05-07-2007 07:57 AM ET (US)	protons in a t butyl group dont signal indicate that they'll have neighbors?
Peter Norris	407 05-07-2007 07:30 AM ET (US)	Anthony - If the game is half as good as 2005 then it'll be superb. 2.30 on May 23rd if you're available - we'll be at University Pizzeria watching on the big screen.
Peter Norris	406 05-07-2007 07:29 AM ET (US)	Jamie - that is a t-butyl group. The protons in question 8 are alpha to C=O and also benzylic so they show further downfield than 2.5 ppm.
Anthony	405 05-06-2007 11:31 PM ET (US)	"Liverpool-Milan final again..." ~ Dr. Norris

Done

What is Organic Chemistry?

The study of the compounds that contain carbon and the reactions of those materials (millions known)

Why a whole year of Organic?

Carbon can bond in multiple ways to form a huge number of different molecules, and these compounds form the basis of many different disciplines, e.g.:

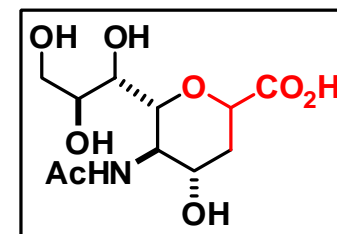
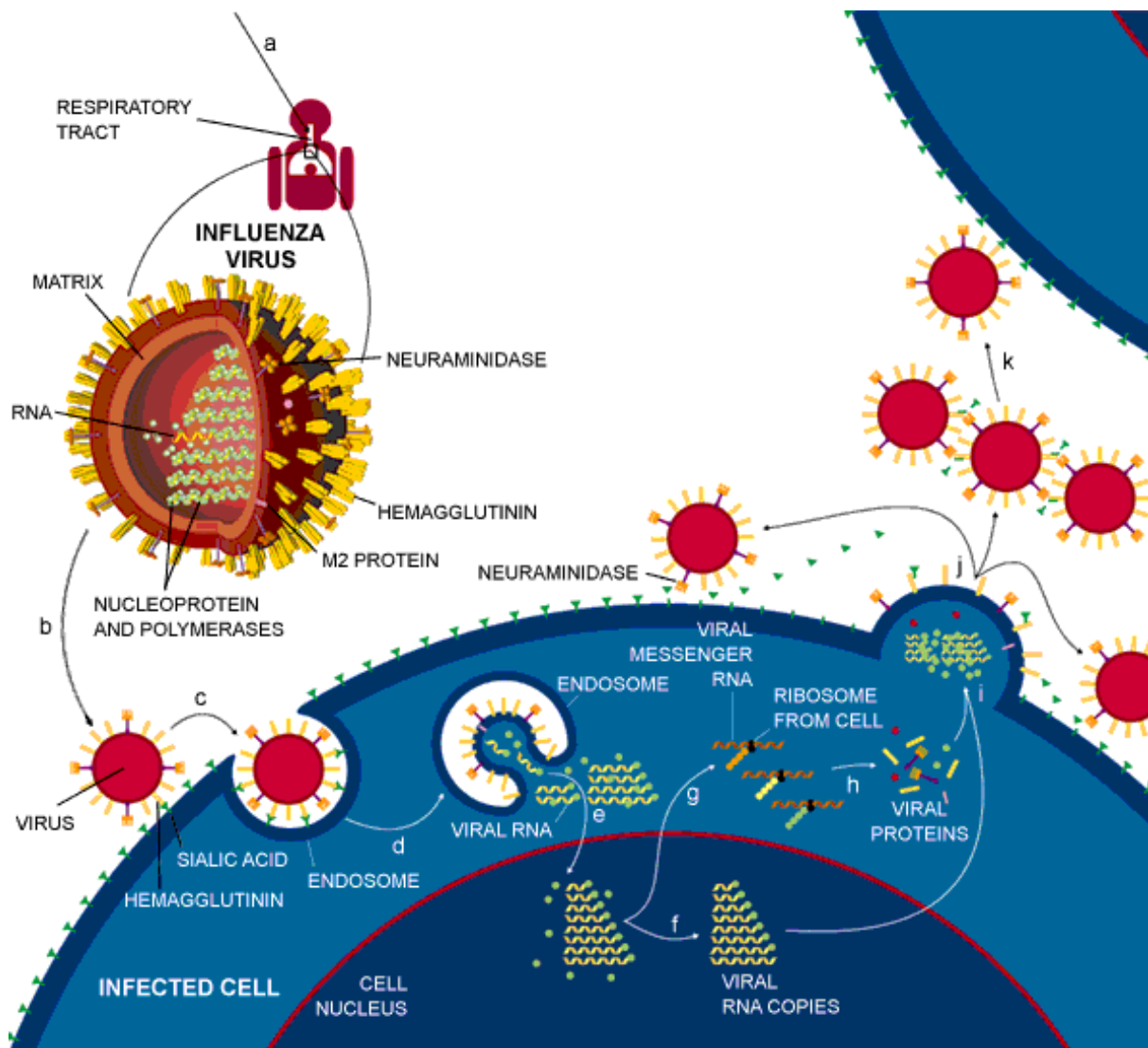
Biology (DNA, proteins, carbohydrates)

Medicine and **Pharmacy** (Aspirin, Taxol, AZT)

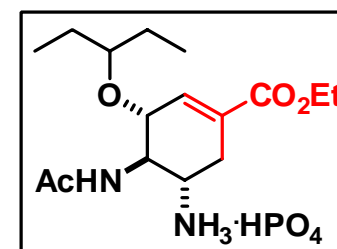
Chemical Engineering (oil, plastics, fine chemicals)

Forensics (Biological materials, chemical tests)

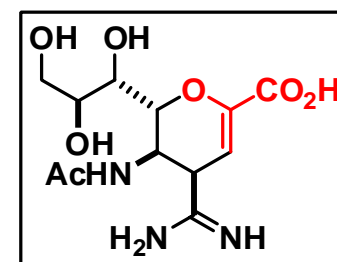
From Organic Chemistry to Biology, Medicine, Pharmacy, etc.



N-acetylneuraminic acid



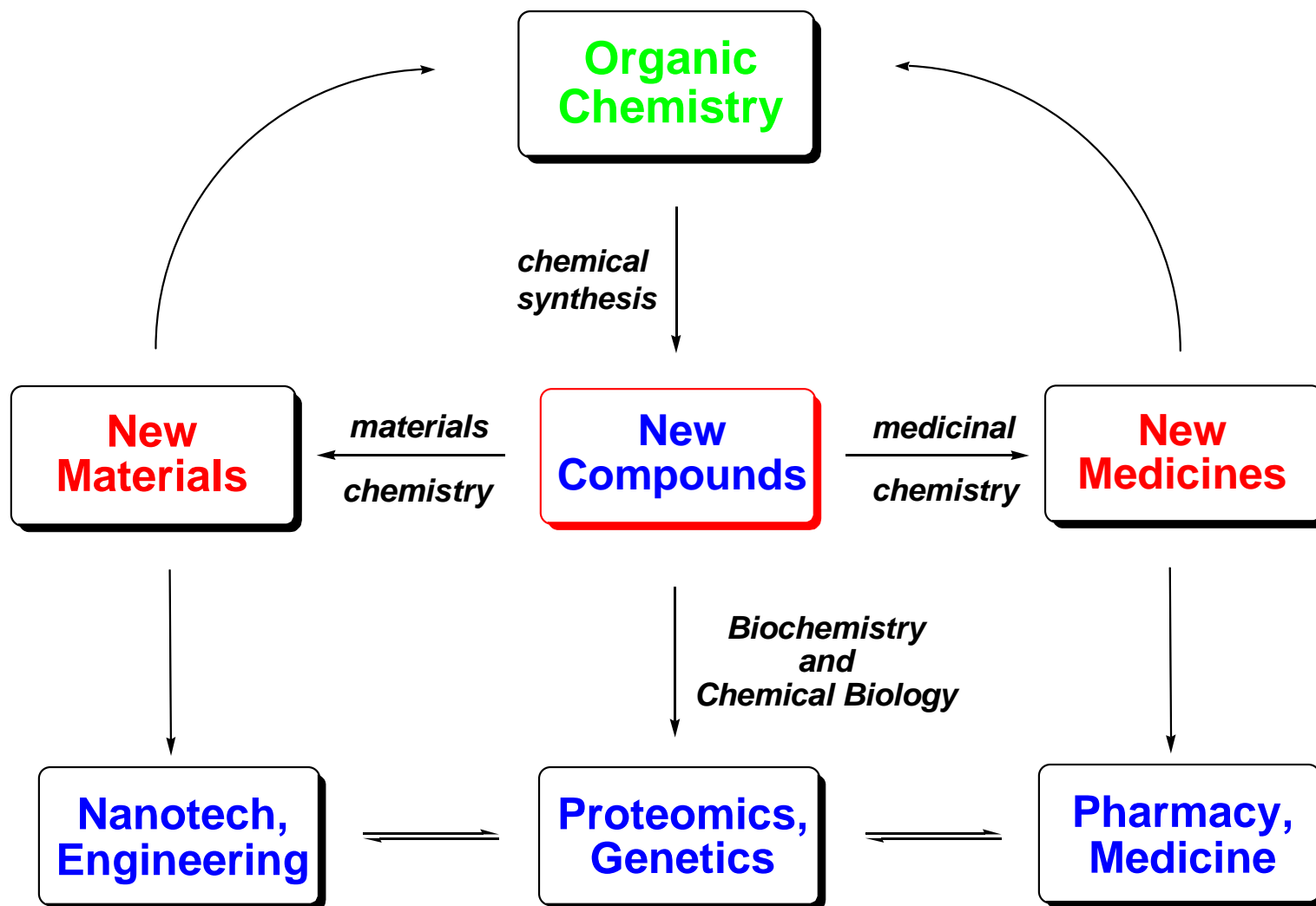
Tamiflu - Giliad/Roche



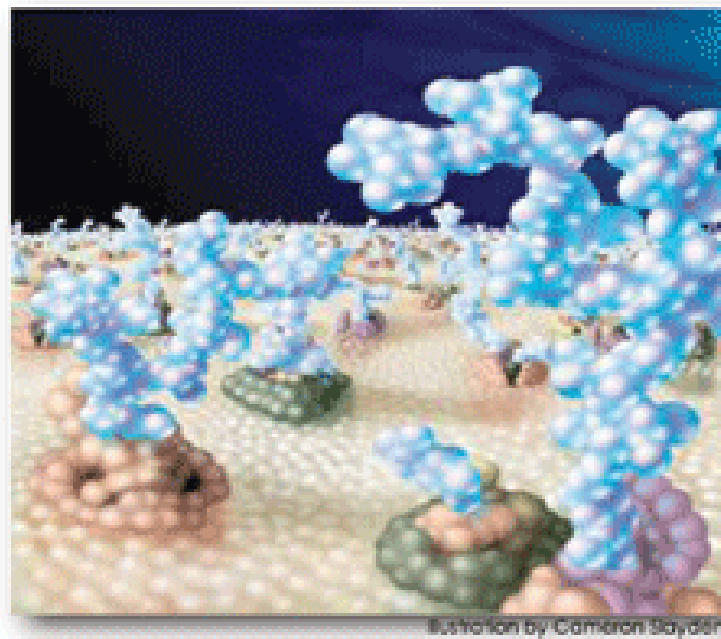
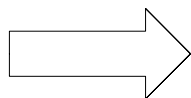
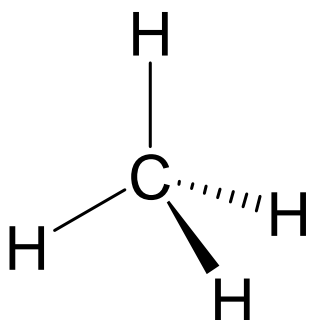
Relenza - GSK

From Scientific American – www.sciam.com

Organic Chemistry – Materials and Uses



Chemistry 3719-3720



~1800 – Organic Chemistry : the chemistry of natural products based on carbon

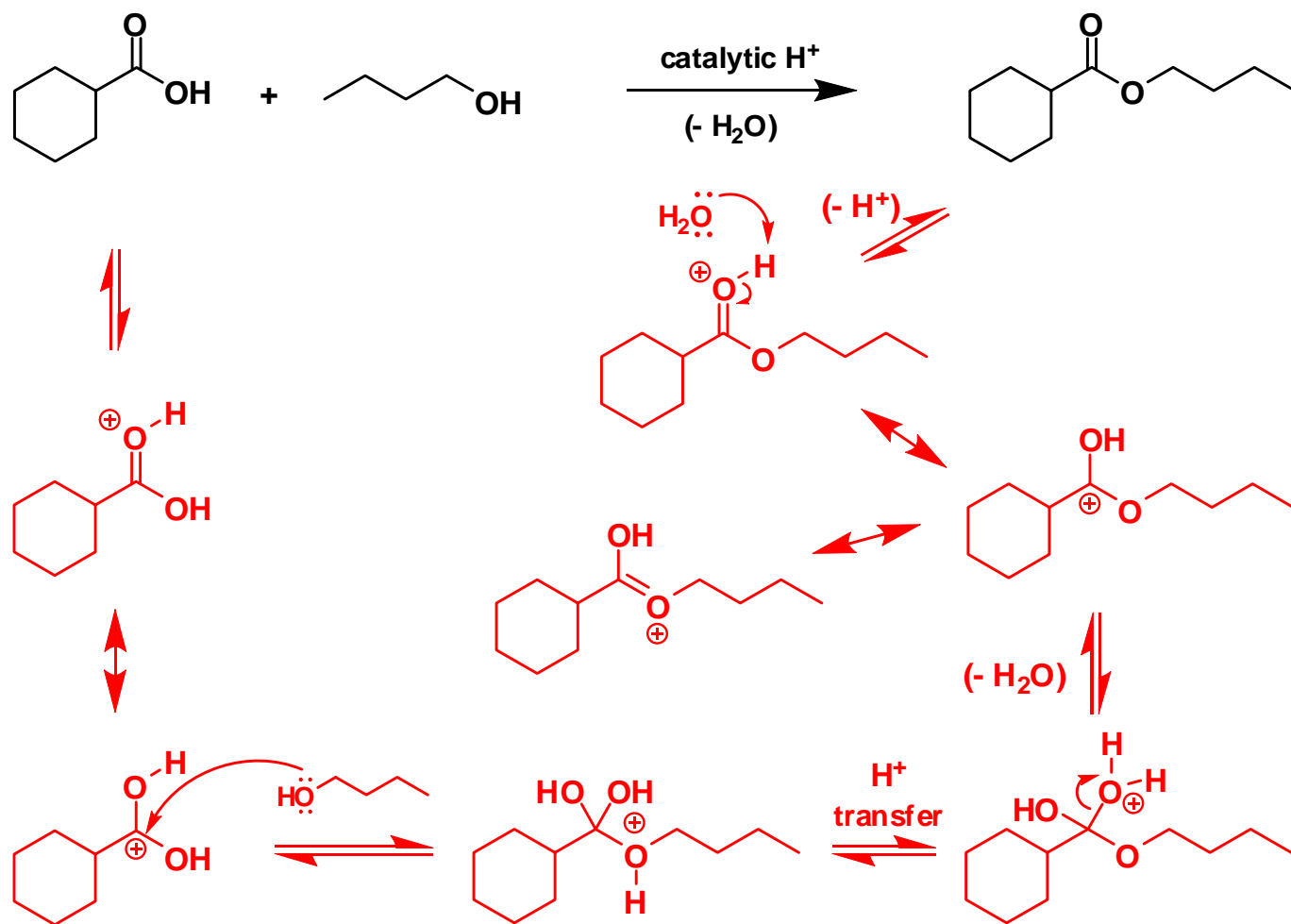
2006 – Organic Chemistry : “molecular engineering”

Timeline

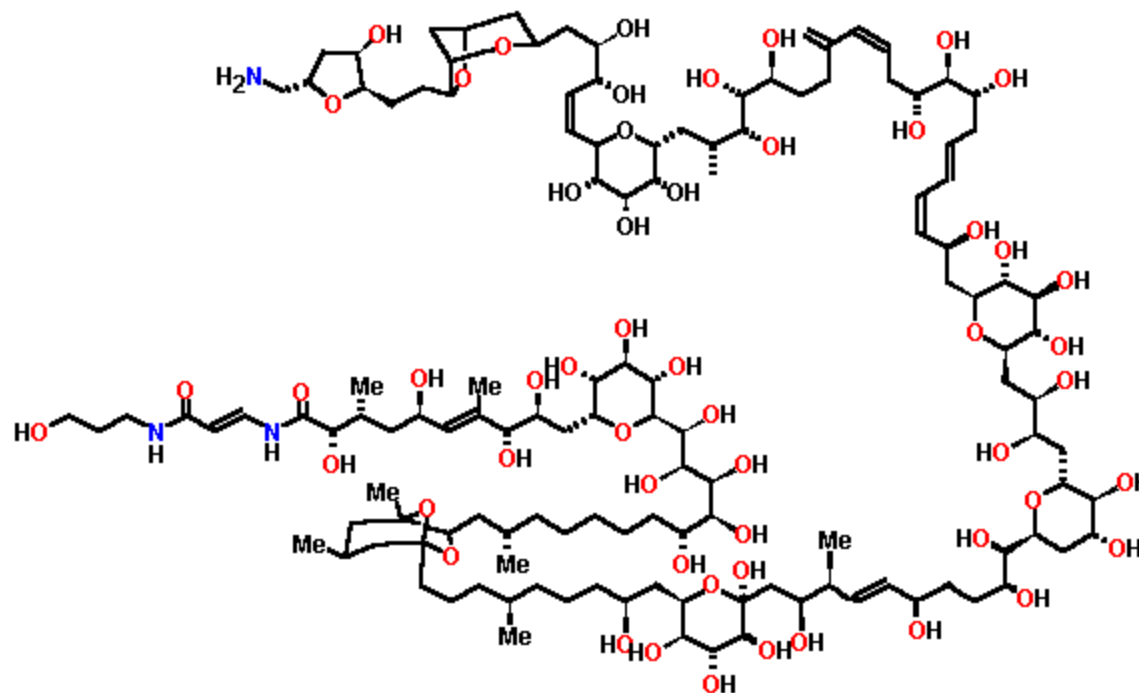
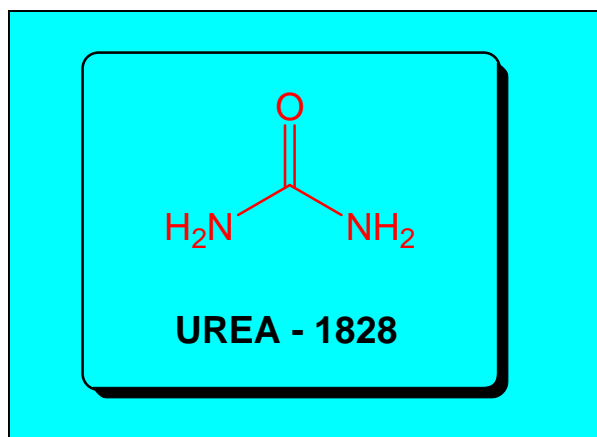
- 1807** Berzelius introduces the term “**Organic Chemistry**” to describe the study of compounds isolated from nature
- 1828** Wöhler makes urea, the first natural organic compound to be synthesized in the laboratory
- 1890** Fischer studies the chemistry of proteins, carbohydrates and the nucleic acids - **Biochemistry**
- 1950** Woodward and Eschenmoser complete the first total synthesis of Vitamin B12. NMR begins to be useful.
- 1990** Kishi, Nicolau, Smith, Schreiber, etc. complete total syntheses of compounds such as Brevetoxin B, Taxol, etc.
- 2000** **Chemical Biology, Molecular Engineering**

Teaching Philosophy: Organic Chemistry as a Language

모든 인간은 태어날 때부터 자유로우며 그 존엄과 권리에 있어 동등하다. 인간은 천부적으로 이성과 양심을 부여받았으며 서로 형제애의 정신으로 행동하여야 한다.

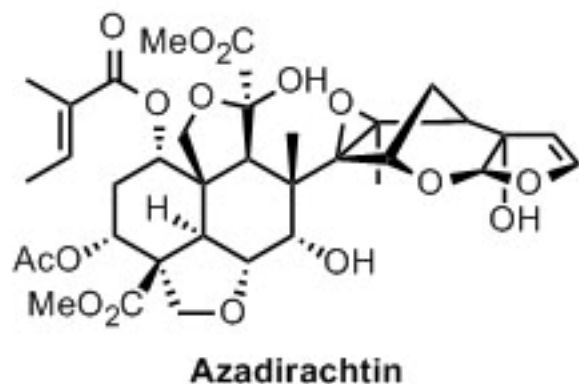


Natural Products Chemistry



Palytoxin

Synthesis: Kishi *et al* J. Amer. Chem. Soc., 111, 7525, **1989**



Ley, Veitch, Beckmann, Burke, Boyer and Maslen. *ACIEE*, August 2007

Carey Chapter 1 - Chemical Bonding

“Structure determines properties”

- Atomic and electronic structure of atoms
- Ionic and covalent bonding
- Electronegativity and polar covalent bonds
- Structures of organic compounds - representations
- Resonance within molecules
- Shapes of molecules
- Molecular orbitals and orbital hybridization

Periodic Table of the Elements

The Modern Periodic Table of the Elements

1												18						
Hydrogen 1 H 1.01 2.1																		Helium 2 He 4.00 ---
2												13	14	15	16	17		
Lithium 3 Li 6.94 1.0	Beryllium 4 Be 9.01 1.5											Boron 5 B 10.81 2.0	Carbon 6 C 12.01 2.5	Nitrogen 7 N 14.01 3.0	Oxygen 8 O 16.00 3.5	Fluorine 9 F 19.00 4.0	Neon 10 Ne 20.18 ---	
Sodium 11 Na 22.99 0.9	Magnesium 12 Mg 24.31 1.2											Aluminum 13 Al 26.98 1.5	Silicon 14 Si 28.09 1.8	Phosphorus 15 P 30.97 2.1	Sulfur 16 S 32.07 2.5	Chlorine 17 Cl 35.45 3.0	Argon 18 Ar 39.95 ---	
Potassium 19 K 39.10 0.8	Calcium 20 Ca 40.08 1.0	Scandium 21 Sc 44.96 1.3	Titanium 22 Ti 47.88 1.5	Vanadium 23 V 50.94 1.6	Chromium 24 Cr 52.00 1.6	Manganese 25 Mn 54.94 1.5	Iron 26 Fe 55.85 1.8	Cobalt 27 Co 58.93 1.8	Nickel 28 Ni 58.69 1.8	Copper 29 Cu 63.55 1.9	Zinc 30 Zn 65.39 1.6	Gallium 31 Ga 69.72 1.6	Germanium 32 Ge 72.61 1.8	Arsenic 33 As 74.92 2.0	Selenium 34 Se 78.96 2.4	Bromine 35 Br 79.90 2.8	Krypton 36 Kr 83.80 3.0	
Rubidium 37 Rb 85.47 0.8	Strontium 38 Sr 87.62 1.0	Yttrium 39 Y 88.91 1.2	Zirconium 40 Zr 91.22 1.4	Niobium 41 Nb 92.91 1.6	Molybdenum 42 Mo 95.94 1.8	Technetium 43 Tc (98) 1.9	Ruthenium 44 Ru 101.07 2.2	Rhodium 45 Rh 102.91 2.2	Palladium 46 Pd 106.42 2.2	Silver 47 Ag 107.87 1.9	Cadmium 48 Cd 112.41 1.7	Indium 49 In 114.82 1.7	Tin 50 Sn 118.71 1.8	Antimony 51 Sb 121.76 1.9	Tellurium 52 Te 127.60 2.1	Iodine 53 I 126.90 2.5	Xenon 54 Xe 131.29 2.6	
Cesium 55 Cs 132.91 0.7	Barium 56 Ba 137.33 0.9	57-70 *	Lutetium 71 Lu 174.97 1.1	Hafnium 72 Hf 178.49 1.3	Tantalum 73 Ta 180.95 1.5	Tungsten 74 W 183.84 1.7	Rhenium 75 Re 186.21 1.9	Osmium 76 Os 190.23 2.2	Iridium 77 Ir 192.22 2.2	Platinum 78 Pt 195.08 2.2	Gold 79 Au 196.97 2.4	Mercury 80 Hg 200.59 1.9	Thallium 81 Tl 204.38 1.8	Lead 82 Pb 207.20 1.8	Bismuth 83 Bi 208.98 1.9	Polonium 84 Po (209) 2.0	Astatine 85 At (210) 2.2	Radon 86 Rn (222) 2.4
Francium 87 Fr (223) 0.7	Radium 88 Ra (226) 0.9	89-102 **	Lanthanum 103 Lr (262) ---	Rutherfordium 104 Rf (261) ---	Dubnium 105 Db (262) ---	Seaborgium 106 Sg (263) ---	Bohrium 107 Bh (262) ---	Hassium 108 Hs (265) ---	Mt 109 Mt (266) ---	Ununium 110 Uun (271) ---	Ununium 111 Uuu (272) ---	Ununium 112 Uub (277) ---	Ununquadium 114 Uuq (289) ---					

Average relative masses are 2001 values, rounded to two decimal places.

All average masses are to be treated as measured quantities, and subject to significant figure rules. Do not round them further when performing calculations.

Element name — Mercury
Atomic # — 80
Symbol — **Hg**
Avg. Mass — 200.59
Electronegativity — 1.9

*lanthanides

Lanthanum 57 La 138.91 1.1	Cerium 58 Ce 140.12 1.1	Praseodymium 59 Pr 140.91 1.1	Neodymium 60 Nd 144.24 1.1	Promethium 61 Pm (145) 1.1	Samarium 62 Sm 150.36 1.2	Europium 63 Eu 151.97 1.1	Gadolinium 64 Gd 157.25 1.2	Terbium 65 Tb 158.93 1.1	Dysprosium 66 Dy 162.50 1.2	Holmium 67 Ho 164.93 1.2	Erbium 68 Er 167.26 1.2	Thulium 69 Tm 168.93 1.3	Ytterbium 70 Yb 173.04 1.1
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**actinides

Actinium 89 Ac (227) 1.1	Thorium 90 Th 232.04 1.3	Protactinium 91 Pa 231.04 1.5	Uranium 92 U 238.03 1.4	Neptunium 93 Np (237) 1.4	Plutonium 94 Pu (244) 1.3	Americium 95 Am (243) 1.3	Curium 96 Cm (247) 1.3	Berkelium 97 Bk (247) 1.3	Californium 98 Cf (251) 1.3	Einsteinium 99 Es (252) 1.3	Fermium 100 Fm (257) 1.3	Mendelevium 101 Md (258) 1.3	Nobelium 102 No (259) 1.3
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1.1 Atoms, electrons, and orbitals

Probability distribution for an electron

Figure 1.1

Boundary surfaces of a 1s and 2s orbital

Figure 1.2

Boundary surfaces of the 2p orbitals

Figure 1.3

Electronic Configurations of Atoms

Electronic Structure of Atoms

<u>Atom</u>	<u>Atomic No.</u>	<u>Electronic Structure</u>
H	1	$1s^1$
He	2	$1s^2$
Li	3	$1s^2 2s^1$
Be	4	$1s^2 2s^2$
B	5	$1s^2 2s^2 2p_x^1$
C	6	$1s^2 2s^2 2p_x^1 2p_y^1$
N	7	$1s^2 2s^2 2p_x^1 2p_y^1 2p_z^1$
O	8	$1s^2 2s^2 2p_x^2 2p_y^1 2p_z^1$

1.1 General Concepts

- Orbitals higher in energy further they are from nucleus.
- Designated by principal quantum number (1, 2, 3, etc.).
- Degenerate orbitals (same energy) fill up singly before they double up (Aufbau).
- Maximum of two electrons per orbital, each having opposite spin (Pauli exclusion principle).
- Impossible to know both the speed and location of an electron at the same time (Heisenberg uncertainty).